A graph of activity and activity

Description automatically generated

Figure 17 - The estimated activities based on two different calculations. Using all COOH binding energies (blue) and using only the COOH binding energies from on-top sites, which does not have a neighbouring hollow site, where H has a negative binding energy. These sites are assumed to be CO-poisoned due to the COOH+H disproportionation reaction, and are counted as a 0 in the activity sum.

A chart of content

Description automatically generated

Figure 19 – Ternary activity plot with 5% molar fractions of platinum, silver and gold. The highest activity found was 5.2\*10e-6 at the composition Pt0.8Ag0.2 at the potential (eU) of 0.07 V vs RHE.

In figure 17, the metals platinum, silver and gold has been chosen, with all 5% molar fractions

A pyramid of content

Description automatically generated with medium confidence

Figure 20 - figure

Fedt

A pyramid of content

Description automatically generated with medium confidence

Figure 21 - Ternary activity plot based on 5% molar fractions of palladium, gold, and silver. The optimal composition when restricted to the three selected metals is Pd0.9Au0.1 with a comparably low activity of 2.4\*10e-10 at a potential (eU) of 0.2 V vs RHE.

Tekst

**DFT calculations**

***Slabs for testing CO-OH slide reaction***

what was the conclusion on this? Seemed unreasonable?

I guess basically the story is: FA has a high theoretical efficiency, but is plagued by very low activity until the voltage has been increased to clean the surface, pointing towards a

Single-site anode catalysts have had success (refx3), so this points towards a hypothesis that is xx, hence it should be investigated how and why and if it can be replicated theoretically and computationally to learn more about it and perhaps peep into optimizations.

***Bayesian optimization***

Math explanation, practical – Explain in introduction or methods?

In figure x, two activity estimation routines are compared on varying potentials (eU) for the same simulated HEA surface. The simplest routine, shown in blue, involves just applying equation x to all binding energies of COOH on on-top sites of the simulated surface. This activity increases as the potential (eU) decreases from 0.2 V and achieves the maximum activity at the optimal potential calculated to -0.17 V. However, formic acid oxidation is assumed to be held back by CO oxidation, so the activity estimation function needs to account for that in order to accurately predict FAOR performance. This routine is shown in orange, and as expected shows an increasing activity when the potential decreases from 0.2 V, but then is decreases by hydrogen UPD, which starts poisoning on-top sites around 0.1 V and has poisoned every site around 0.16 V. A maximum is found in the middle, typically between at small positive potentials around 0.07 V, depending on the surfaces ability to deter H adsorbing to neighbouring on-top sites. A theoretical surface with large positive binding energies for H in hollow sites and low COOH binding energies on on-top sites would be able to be most active at a lower potential, which would improve the efficiency.

A brute-force search approach has been used on the stoichiometries / molar fractions to find the most active composition of the surface. Molar fractions of 5% intervals was made for all five metals, leading to 10626 compositions, and from each composition a 200 by 200 by 3 surface was simulated, the binding energies of COOH and H predicted by the appropriate model and the activity estimated.

Based on the COOH binding energies as shown in figure 8, it’s known that either platinum or palladium is needed for formic acid oxidation, but pure platinum or palladium surfaces might not be optimal anode catalysts due to CO poisoning. Hence, examining the impact of silver, gold, and copper along with platinum and/or palladium on the formic acid oxidation activity is of interest.

When testing all 5% molar fraction the highest activity is found at this stoichiometry at this potential.

With ternary activity plots, plotting the activity of all possible molar fractions of a trimetal is possible.

The hunt is therefore on for an atom on the periodic table, which fulfils the criteria for an optimal catalyst.

Scaling relations might be a hindrance

COOH needs to bind way stronger, but that is correlated with a stronger H binding in general. Can this be circumvented with single-sites or will H just move on-top and block?

The optimal anode potential sets some pretty rough criteria. Also, the activity increases at higher potentials and the efficiency per HCOOH increases with the open-circuit voltage as

De forskellige potentialer vs RHE

CO oxidation and CO creation.

Note: Nifty that I have neighbour DFT data that might be able to show the same thing.

Note: Go into specifics with different reaction pathways? Mention the dentates? Or just show the free energy diagram later and say it doesn’t look feasible? Based on DFT data the energy increases.

Show free energy diagrams – optimal binding energy and perhaps one with voltage that makes them equal in energy? And one with the energies from

Måske vis, at man kan bruge Pt som reference, og så notere bindingsenergier som “den binder så godt her ift. Platin”, og det har man for OH og det dobbelte for O, men der er ikke for COOH og dentaterne. (Eller ligegyldigt måske? Jeg burger det jo ikke)

Kohn-Sham DFT applied to calculate the binding energies of various adsorbates.

Hohenberg-Kohn proved, that electron densities can perfectly describe a quantum system. Two theorems. 1: Unique one-to-one mapping between ground state energy of system and its electron density. 2: A variational principle holds for electron densities; the electron density can be found by minimizing the expectation value of the energy (Jacks sætning) (Det hele skal omskrives en gang ekstra, selvom intet er kopieret direkte)

**Plotting GH vs GCOOH**

Plotting them H vs COOH

Using energy prediction models on simulated surfaces

The optimal achievable U\_anode might be found in the region just above the potentials, where both H and COOH is present, even though lower COOH binding energies would be more efficient in terms

Then, the challenge of finding the optimal

Describe the criteria used to judge stochs through their binding energies

Swim ring combinations binding energy plots

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| Fuel | Gravimetric energy density (kWh/kg) | Volumetric energy density (kWh/L) | Max theoretical efficiency | Reaction steps | Price (US$/kg, retail) | Operating temperature (°C) | Storage vessel | Storage pressure |
| FA |  |  |  |  |  |  |  |  |
| H2 |  |  |  |  |  |  |  |  |
| CH3OH |  |  |  |  |  |  |  |  |

The electrical energy released by FAOR per electron can be estimated with the free-energy diagram of the direct pathway, as seen in figure x.

FOAR highest U vs CHE: negative

OER highest U vs CHE: positive

The difference is the V\_OC

#If formate binds too weakly, the potential can’t be lowered much, before the binding reaction is no longer favourable, leading to a small open-circuit voltage. If formate binds too strongly, what happens? (Tjek ask Jack).

**Notes for formic acid:**

Already close to CO2 in structure, compare to other species like methanol, ethanol. More effective with fever electron transfers and steps. Plot similar to the one from Jack with energy efficiencies

Methanol has a CO base, while formic acid has a OCO base. Methanol has 4 H. So the methanol oxidation requires a pretty large overpotential, which FAOR doesn’t need. This means the potential optimal decent working potential is lower for FAOR.

# How much energy comes out? Per electron? To calculate the maximum electrical energy

Make a simple figure that shows both methanol and formic acid with the backbone highlighted.

Kilder til HEA:

1. [15]  Z. Y. Lv, X. J. Liu, B. Jia, H. Wang, Y. Wu, and Z. P. Lu. Development of a novel high-entropy alloy with eminent efficiency of degrading azo dye solutions. Scientific Reports, 6(34213), 2016. https://doi.org/10.1038/ srep34213.
2. [16]  Chih-Fang Tsai, Pu-Wei Wu, Pang Lin, Cheun-Guang Chao, and Kung-Yu Yeh. Sputter Deposition of Multi-Element Nanoparticles as Electrocatalysts for Methanol Oxidation. Japanese Journal of Applied Physics, 47(7R):5755, 2008. https://doi.org/10.1143/JJAP.47.5755.

Her er Kohn-Sham-Hohenberg-referencerne:

1. [20]  P. Hohenberg and W. Kohn. Inhomogeneous electron gas. Physical Review, 136(3B):B864–B871, 1964. https://doi.org/10.1103/PhysRev.136.B864.
2. [21]  W. Kohn and L. J. Sham. Self-consistent equations including exchange and correlation effects. Physical Review, 140(4A):A1133–A1138, 1965. https: //doi.org/10.1103/PhysRev.140.A1133.